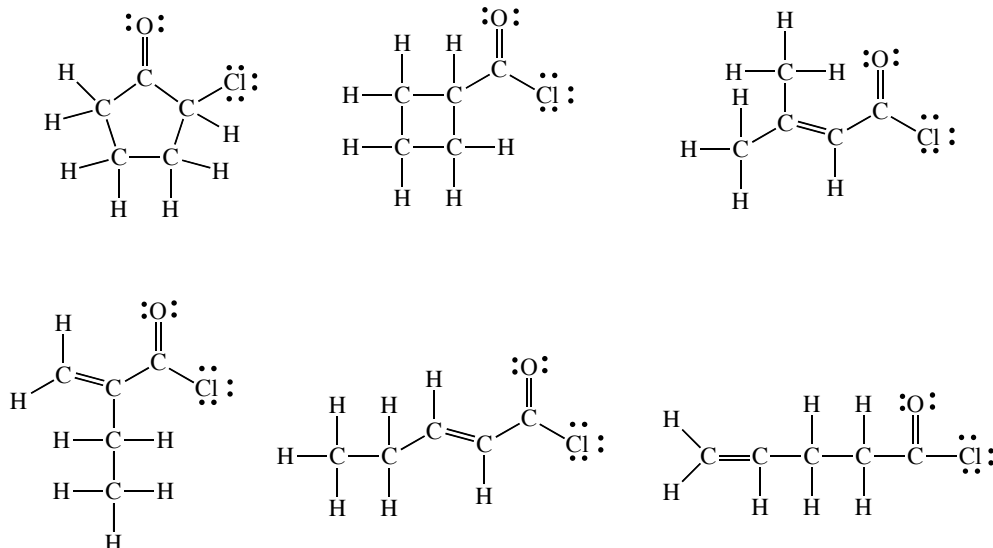


Name: \_\_\_\_\_ ANSWER KEY \_\_\_\_\_

## Chemistry 250A -- September 10, 2009 -- HW/Quiz #1 Answer Key

To receive full credit *show all lone pair electrons and non-zero formal charges.*

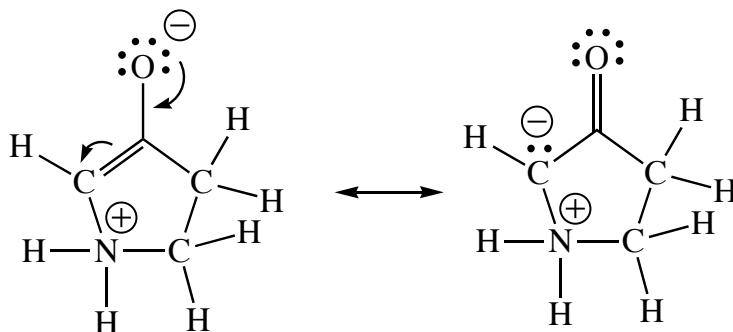
1. (5 pts) Aldrich Chemical Company sells six compounds with the formula  $C_5H_7ClO$ . Draw Lewis structures for 3 *stable* molecules (constitutional isomers) having this molecular formula.



**Here are six that Aldrich sells. There are many other possible isomers.**

2. (5 pts)

- Complete the Lewis structure shown below by indicating all non-zero formal charges.
- Draw an additional important resonance form for this molecule. (Hint: All atoms should have filled shells.)
- Circle the resonance form that is most important.



Most important resonance form.  
(Negative charge on more electronegative atom.)